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spending more cash.
nevertheless when?
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PRACTICE PROBLEM 2

(page 324) 2. Sulfur

burns in oxygen to
form sulfur dioxide.

Write a skeleton
equation for this
chemical reaction.

Include appropriate
symbols from table

11.1. Analyze Step

1. Identify the relevant
concepts. Write the

formula for each
reactant and each

product. Include the

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common

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SECTION 11.1

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**REACTIONS (pages
321-329)**

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A representation of a
chemical reaction with
reactants on the left,
products on the right,
and an arrow
separating the two.

Small whole number
ratios that are placed

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in front of the formulas
in an equation in order
to balance it.

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chemical reaction, the

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reactants are written on the left and the products on the right. The arrow that separates them is called yield.

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The law of
conservation of mass
states that the mass of

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the reactants will always equal the mass of the products. In result the number of atoms of one element on the reactants side should be identical to the number of atoms of the same element on the products side.

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Include the common

SECTION 11.1

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(page 324) 2. Sulfur
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Write a skeleton
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Include appropriate
symbols from table
11.1.

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KEY PDF

Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed, and percent yield based on theoretical maximum. ... 15.2

CHAPTER 11:
STOICHIOMETRY. MOLE
TO MOLE RATIO. When
nitrogen and hydrogen
gas are heated under
the correct conditions,

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ammonia gas (NH_3) is
formed.

Reading Answer

CHAPTER 11:

STOICHIOMETRY

Chapter 11 Guided

Notes. Name: _____

- 11.1 - Describing
Chemical Reactions. 1.
What two substances
reacted in the
Hindenburg crash. 2. In
general, what happens
in a chemical reaction.
3. Look at Figure 11.1.
What are some of the
evidences of, chemical

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change (a.k.a.

chemical reaction) 4.

Give an example of a
word equation for a

chemical reaction (you

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Stoichiometry

Problems. When we carry out a reaction in either an industrial setting or a laboratory, it is easier to work with masses of substances than with the numbers

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of molecules or moles.
We will first present the method used in most other books for converting from the mass of any reactant or product to the mass of any other reactant or product using a balanced chemical equation outlined ...

Chapter 11.4:
Stoichiometry -
Chemistry
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oxidation–reduction reactions A chemical reaction that exhibits a change in the oxidation states of one or more elements in the reactants that has the general form oxidant + reductant \rightarrow reduced oxidant + oxidized reductant. The general forms of these five kinds of reactions are

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Key

summarized in Table
11.6.1, along with
examples of each.

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